

T04D02 - Formal Charge

Name

Formal charge is a way to “select” the more important resonance structure.

The formula for formal charge is:

Formal charge = Number of valence electrons – [$\frac{1}{2}$ number of shared electrons + number of unshared valence electrons]

a. Sulfur trioxide

b. Laughing gas, N₂O

c. Cyanate ion, NCO⁻¹

d. Carbon dioxide

e. Carbon disulfide

f. Thiocyanate ion, SCN⁻¹

